

Oxidizing: reactant of reduction agent

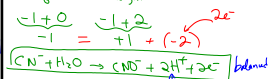
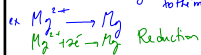
Reducing agent: reactant of oxidation

Full reaction \rightarrow Redox

$\frac{1}{2}$ reaction \rightarrow oxidation ex: $Mg \rightarrow Mg^{2+}$
 \rightarrow reduction ex: $CN^- \rightarrow CNO^-$

Balancing: \checkmark atoms

charge: adding electrons to the more positive side



Condition: acidic $\rightarrow H^+$ add
 basic $\rightarrow OH^-$ add

Switching acid \leftrightarrow base condition

for every H^+ , OH^- is needed
 $H^+ + OH^- \rightarrow H_2O$ (Add "OH" to both sides)

